

WRITING NET IONIC EQUATIONS Pre-AP

Write a balanced net ionic equation for the reactions that occurs in each of the following cases.

Steps for writing net Ionic Equations

1. Write the balanced molecular equation.
2. Write the balanced complete ionic equation (all soluble ions listed).
3. Cross out the spectator ions that are present.
4. Write the "leftovers" as the net ionic equation.

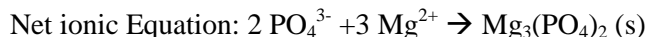
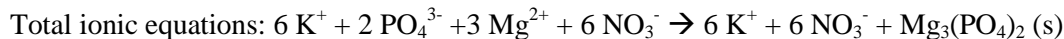
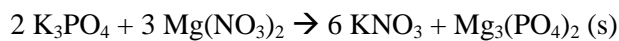
Solubility Rules*

Ion	Solubility	Exceptions
NO_3^-	soluble	none
ClO_4^-	soluble	none
Cl^-	soluble	except Ag^+ , Hg_2^{2+} , Pb^{2+}
I^-	soluble	except Ag^+ , Hg_2^{2+} , Pb^{2+}
SO_4^{2-}	soluble	except Ca^{2+} , Ba^{2+} , Sr^{2+} , Hg^{2+} , Pb^{2+} , Ag^+
CO_3^{2-}	insoluble	except Group IA and NH_4^+
PO_4^{3-}	insoluble	except Group IA and NH_4^+
OH^-	insoluble	except Group IA, Ca^{2+} , Ba^{2+} , Sr^{2+}
S^{2-}	insoluble	except Group IA, IIA and NH_4^+
Na^+	soluble	none
NH_4^+	soluble	none
K^+	soluble	none

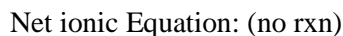
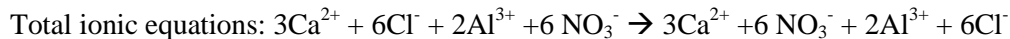
*slightly soluble

Complete on a separate sheet of paper

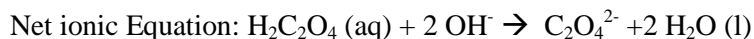
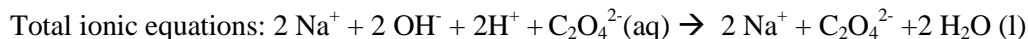
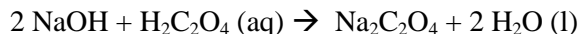
1. potassium phosphate + magnesium nitrate \rightarrow



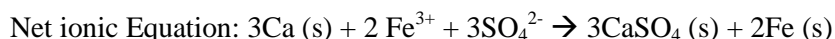
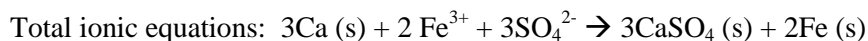
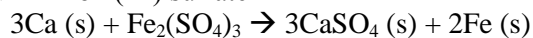
2. calcium chloride + aluminum nitrate \rightarrow

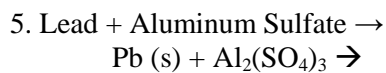


3. sodium hydroxide + oxalic acid \rightarrow



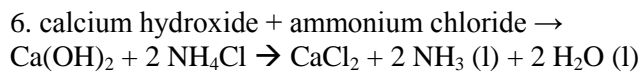
4. calcium + iron (III) sulfate \rightarrow



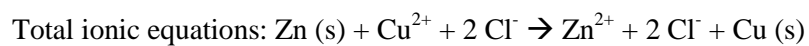
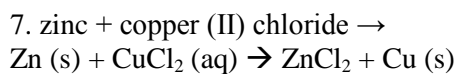
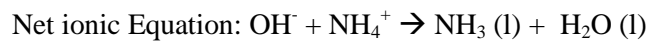


Total ionic equations:

Net ionic Equation:



Total ionic equations:



Net ionic Equation:

Write net ionic equations for the following reactions:

